

**CSSR & SRRM DEGREE & PG COLLEGE  
(Autonomous)**

**Kamalapuram - 516289, Kadapa, A.P.**



Department of Chemistry

Board of Studies

Minutes of Meeting-I

AY: 2025-26

Dated:14/06/2025

## AGENDA

1. To approve syllabus of III semester and IV Semester chemistry (Minor) courses for II B.Sc. Honours (Botany)
2. To approve course outcomes, revised syllabus, Blue Print, Model Question Paper, for
  - (i) Chemistry- Paper-II-Fundamental in Organic Chemistry (Minor Paper) for II Year, Semester - III, B.Sc. Honours (Botany) in the academic year 2025-26
  - (ii) Chemistry-Paper-III-Physical Chemistry-II (States of matter, Phase rule and surface chemistry). (Minor Paper) for II Year, Semester - IV, B.Sc. Honours (Botany) in the academic year 2025-26.
  - (iii) Chemistry-Paper-IV-General and physical chemistry (Minor Paper) for II Year, Semester - IV, B.Sc. Honours (Botany) in the academic year 2025-26
3. To approve the course structure, Continuous Internal Assessment (CIA), Semester End Examination (SEE) Practical examinations Patterns
4. To suggest panel of names for appointment of paper setters and examiners
5. To approve other academic activities of the department.
6. Any discussions with approval of the chair.

**The Modified syllabus by the panel members in the BoS Meeting are:**

Sl. No.	Semester, Group & Title of the Paper	Modifications	Remarks
1	Fundamentals in Organic Chemistry	<b>UNIT-V Orientation of aromatic substitution</b> Removed the topics: nitrile, carbonyl	In this unit syllabus is large.

*Y.v. Lawhmi*

### Course Structure

Year	Semester	Course	Program Code	Title of the Course	No. of Hrs/Week	No. of Credits	IA	EA	TOTAL
II	III	2	BOT305	Fundamentals in Organic Chemistry - (Theory)	3	3	30	70	100
				Fundamentals in Organic Chemistry - (Practical)	2	1	-	50	50
II	IV	3	BOT404	Physical Chemistry-II (States of Matter, Phase Rule & surface Chemistry) -	3	3	30	70	100
				Physical Chemistry-II (States of Matter, Phase Rule & surface Chemistry) - (Practical)	2	1	-	50	50
II	IV	4	BOT405	General & Physical Chemistry (Theory)	3	3	30	70	100
				General & Physical Chemistry (Practical)	2	1	-	50	50

### Continuous Internal Assessment:

In each semester, for every subject there are two Internal Examination with 30 marks each and time duration of 1 Hour. The Thirty marks are divided as:

Sl. No	Name of the Activity	Marks Allotted
1	Internal Examination	20
2	Co Circular Activities: Seminar/Assignment/JAM/Group Discussion	5
3	Extra Circular Activities: NSS/NCC/Sports/Clean & Green Activities/Community Services	5

### Semester End Examinations:

The semester end examination is for 70 marks with the time duration of 3 Hours.

### Practical Examinations:

The Practical examination is for 50 marks with the time duration of 3 Hours

Sl. No	Name of the Activity	Marks Allotted
1	Record	10
2	Viva voice	10
3	Practical	30

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Syllabus

**COURSE OUTCOMES:**

At the end of the course the student will be able to-

CO1: Understand and explain the differential behavior of organic compounds based on fundamental concepts learnt.

CO2: Formulate the mechanism of organic reactions by recalling and correlating the fundamental properties of the reactants involved.

CO3: Learn and identify many organic reaction mechanisms.

CO4: Correlate and describe the stereo-chemical properties of organic compounds and reactions.

CO5: Understand how aromaticity influences the chemical behavior of a compound, such as its reactivity and stability.

**Unit 1. Structural theory in Organic Chemistry (9 h)**

Types of bond fission and organic reagents (Electrophilic, Nucleophilic, and free radical reagents). Reaction intermediates – Carbocations, carbanions & free radicals. Bond polarization: Factors influencing the polarization of covalent bonds, inductive effect - Application of inductive effect (a) Basicity of amines (b) Acidity of carboxylic acids (c) Stability of carbonium ions. Resonance or Mesomeric effect, application to (a) acidity of phenol, and (b) acidity of carboxylic acids. Hyper conjugation and its application to stability of carbonium ions, Free radicals and alkenes.

**Unit II Saturated Hydrocarbons (Alkanes and Cyclo alkanes) (9 h)**

General methods of preparation of alkanes- Wurtz and Wurtz Fittig reaction, Corey House synthesis, physical and chemical properties of alkanes, Conformational analysis of alkanes (Conformations, relative stability and energy diagrams of Ethane, Propane and butane). General molecular formulae of cycloalkanes and relative stability, Baeyer strain theory, Cyclohexane conformations with energy diagram, Conformations of monosubstituted cyclohexane.

**UNIT-III Unsaturated Hydrocarbons (Alkenes and Alkynes) (9 h)**

General methods of preparation, physical and chemical properties, Saytzeff and Hoffmann eliminations (with mechanism), Electrophilic Additions, (H<sub>2</sub>, HX) mechanism (Markonikoff/ Antimarkownikoff addition) with suitable examples-syn and anti-addition;

Addition of X<sub>2</sub>, HX. Oxymercuration demercuration, ozonolysis, hydroxylation, Diels Alder reaction, 1,2- and 1,4-addition reactions in conjugated dienes. Reactions of alkynes; acidity, electrophilic and Nucleophilic additions, hydration to form carbonyl compounds, Alkylation of terminal alkynes.

**UNIT-IV Benzene and its reactivity (9 h)**

Structure of Benzene – Preparation - polymerisation of acetylene and decarboxylation- Properties -mechanism of electrophilic aromatic substitution of Friedel- Craft's alkylation and Acylation. Halogenation and nitration,

**UNIT-V Orientation of aromatic substitution (9 h)**

Concept of aromaticity, Huckel's rule - application to Benzenoid (Benzene, Naphthalene) and Non - Benzenoid compounds (cyclopropenylcation, cyclopentadienyl anion and tropylium cation) Orientation of aromatic substitution - ortho, Para and meta directing groups. Ring activating and deactivating groups with examples (Electronic interpretation of various groups like NO<sub>2</sub> and Phenolic). Orientation of (i) Amino, methoxy and methyl groups (ii) Carboxy, nitro and sulphonic acid groups (iii) Halogens.

**List of Reference Books**

1. Morrison, R. N. & Boyd, R. N. Organic Chemistry, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education).
2. Finar, I. L. Organic Chemistry (Volume 1), Dorling Kindersley (India) Pvt. Ltd. (Pearson Education).
3. Guide book to Mechanism in Organic Chemistry by Peter Sykes 6th edition.

Y. V. Lakshmi

### III -SEMESTER

Title of the Course: Organic Qualitative analysis (Practical)

Credits: 01

#### Course outcomes:

At the end of the course, the student will be able to;

1. Use glassware, equipment and chemicals and follow experimental procedures in the laboratory
2. Determine melting and boiling points of organic compounds
3. Understand the application of concepts of different organic reactions studied in theory part of organic chemistry

#### Syllabus:

Analysis of an organic compound through systematic qualitative procedure for functional group identification including the determination of melting point and boiling point with suitable derivatives. Alcohols, Phenols, Aldehydes, Ketones, Carboxylic acids, Aromatic primary amines, amides and simple sugars.

- Co-curricular activities and Assessment Methods
- Continuous Evaluation: Monitoring the progress of student's learning.
- Class Tests, Work sheets and Quizzes

#### Reference books:

- 1) Vogel A.I .Practical Organic Chemistry, Longman Group Ltd.
- 2) Bansal R.K. Laboratory Manual of Organic Chemistry, Wiley-Eastern.
- 3) Ahluwalia V. K. and Agarwal R. Comprehensive Practical Organic Chemistry, University press.

Y.v. Lavshmi

II B.Sc. Honours (Botany)  
Chemistry (Minor)

**BLUE PRINT**  
for All the Courses in every Semester

Duration: 3 Hrs.

Total Marks: 70

**SECTION-A**

Answer any Five questions. Each question carries 4 Marks.

5×4 = 20 M

Topics	Questions Given	Allotted Marks
UNIT-1	1	4
	2	
UNIT-2	3	4
	4	
UNIT-3	5	4
	6	
UNIT-4	7	4
	8	
UNIT-5	9	4
	10	

**SECTION-B**

Answer ALL the Questions. Each question carries 10 MARKS.

5×10 = 50 M

Topics	Questions Given	No. of Questions to be Answered	Allotted Marks
UNIT-1	11	1	10
	12		
UNIT-2	13	1	10
	14		
UNIT-3	15	1	10
	16		
UNIT-4	17	1	10
	18		
UNIT-5	19	1	10
	20		

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II B.Sc. Honours (Botany)  
Semester-III\_Minor- Chemistry

Title: Fundamentals in Organic Chemistry

MODEL QUESTION PAPER

Time: 3 Hours

Max. Marks: 70

Section – A

Answer any FIVE questions

(5 x 4 = 20 Marks)

- 1) Write types of organic reagents with a suitable example.
- 2) Discuss Acidic nature of phenol with mesomeric effect.
- 3) Write any two preparation methods of cycloalkanes.
- 4) Write short note on
  - (a) Corey-House synthesis
  - (b) Wurtz reaction
- 5) Write short note on Anti-Markonikoff rule.
- 6) Give an account of Acidic nature of terminal alkynes.
- 7) Write the mechanism for Halogenation of Benzene.
- 8) Write the mechanism for Nitration of Benzene.
- 9) Discuss effect of Ortho directing group on electrophilic substitution of Benzene.
- 10) Discuss effect of Para directing group on electrophilic substitution of Benzene.

Section – B

Answer ALL the Questions.

5×10 = 50 M

11. What is Inductive effect? Write any two application of Inductive effect

(OR)

12. Explain Hyper conjugation and its application to stability of carbonium ions.

13. Explain Baeyer's Strain Theory.

(OR)

14. Write conformational isomers and energy diagrams of Propane and n-butane.

15. Explain reactions of Alkenes with

(a) Bromine water

(b) Alkaline  $\text{KMnO}_4$

(OR)

16. Explain 1,2- and 1,4-addition reactions in 1,3-Butadiene

17. Write Friedel Craft Alkylation and Acylation reaction

(OR)

18. Discuss the structure of Benzene with M.O Diagram

19. Explain Ring activating group and deactivating groups with an example.

Or

20. What is Huckel's rule? Write applications to Benzenoid and Non-Benzenoid compounds.

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IV - SEMESTER  
Title: PHYSICAL CHEMISTRY- II  
(States of Matter, Phase Rule & Surface Chemistry)  
Credits: 03

**Course outcomes:**

- CO1: Explain the difference between solids liquids and gases in terms of intermolecular interactions.
- CO2: Differentiate ideal and real gases.
- CO3: Discuss the basic concepts of two component systems
- CO4: Apply the concepts of adsorption.
- CO5: Understand the basic concepts of crystallography

**Syllabus:**

**UNIT I - GASEOUS STATE (9 Hours)**

Postulates of Kinetic theory of Gases (exclude derivation) – deduction of gas laws from Kinetic gas equation- Vander Waal's equation of state. Andrew's isotherms of carbon dioxide, continuity of state. Critical phenomena. Relationship between critical constants and Vander Waal's constants. Law of corresponding states. Joule- Thomson effect. Inversion Temperature.

**UNIT II – LIQUID STATE (9 Hours)**

Physical properties of liquids; vapour pressure, surface tension and coefficient of viscosity, and their determination. Effect of addition of various solutes on surface tension and viscosity. Temperature variation of viscosity of liquids and comparison with that of gases. Qualitative discussion of structure of water. Liquid crystals, mesomorphic state. Differences between liquid crystal and solid/liquid. Classification of liquid crystals into Smectic and Nematic. Application of liquid crystals as LCD devices.

**UNIT-III - Solid State (9 Hours)**

Symmetry in crystals. Law of constancy of interfacial angles. The law of rationality of indices. Miller indices, Definition of lattice point, space lattice, unit cell. Bravais lattices and crystal systems. X-ray diffraction and crystal structure. Bragg's law and its derivation. Powder method. Defects in crystals. Stoichiometric and non-Stoichiometric defects.

**Unit IV - Phase Rule (9 Hours)**

The Concept of phase, components, degrees of freedom. Gibbs phase rule. Phase diagram of one component system – water system, Study of Phase diagrams of Simple eutectic systems i) Pb-Ag system, desilverisation of lead ii) NaCl-Water system, Congruent and incongruent melting point- Definition and examples for systems having congruent and incongruent melting point, freezing mixtures.

**Unit V Surface Chemistry(9 Hours)**

Definition and classification of Colloids- Coagulation of colloids- Hardy-Schulze rule. Stability of colloids, Protection of Colloids, Gold number. Adsorption - Physical and chemical adsorption, Freundlich and Langmuir adsorption isotherm, applications of adsorption.

**III. List of Reference Books:**

- 1) Solid State Chemistry and its applications by Anthony R. West
- 2) Text book of physical chemistry by K L Kapoor Vol.1
- 3) Text book of physical chemistry by S Glasstone
- 4) Advanced physical chemistry by Bahl and Tuli.

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IV - SEMESTER  
Title: Organic Preparations (Practical)  
Credits: 01

**Course outcomes:**

At the end of the course, the student will be able to:

- 1) Use glassware, equipment and chemicals and follow experimental procedures in the laboratory
- 2) Apply concepts of surface chemistry in experiments.
- 3) Be familiar with the concepts & practical applications of Surface tension and viscosity of liquids.

**Physical Chemistry Practical Syllabus:**

1. Determination of surface tension of liquid by drop count method
2. Determination of surface tension of liquid by drop weight method
3. Determination of surface tension of mixture (liquid + detergent) using stalagmometer.
4. Determination of coefficient of viscosity of an organic liquid.
5. Determination of composition of a glycerol in glycerol + water mixture using Viscometer.
6. Adsorption of acetic acid on animal charcoal, verification of Freundlich isotherm.

**Co-curricular activities and Assessment Methods:**

- 1) Continuous Evaluation: Monitoring the progress of student's learning
- 2) Class Tests, Worksheets and Quizzes
- 3) Presentations, Projects and Assignments and Group Discussions: Enhances

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IV - SEMESTER  
GENERAL AND PHYSICAL CHEMISTRY  
Credits: 03

**Course outcomes:**

1. CO1: Correlate and describe the stereo chemical properties of organic compounds.
2. CO2: Explain the biological significance of various elements present in the human body.
3. CO3: Apply the concepts of ionic equilibrium for the qualitative and quantitative analysis.
4. CO4: Determine the order of a chemical reaction.
5. CO5: Describe the basic concepts of enzyme catalysis.

**Syllabus**

**UNIT-I Stereo chemistry of carbon compounds (9 h)**

Molecular representations - Wedge, Fischer, Newman and Saw-Horse formulae. Optical isomerism: Optical activity- wave nature of light, plane polarized light, optical Rotation and specific rotation. Chiral molecules- definition and criteria (Symmetry elements)-Definition of enantiomers and diastereomers – Explanation of optical isomerism with Examples- Glyceraldehyde, Lactic acid, Alanine, Tartaric acid, 2,3-dibromopentane.

**Unit II Bioinorganic Chemistry (9 h)**

Metal ions present in biological systems, classification of elements according to their action in biological system. Geochemical effect on the distribution of metals, Na / K- pump, Carbonic anhydrase and Carboxy peptidase. Excess and deficiency of some trace metals. Toxicity of metal ions (Hg, Pb, Cd and As), reasons for toxicity, Use of chelating agents in Medicine, Cisplatin as an anti-cancer drug. Iron and its application in bio-systems, Hemoglobin-transfer of oxygen, Myoglobin-Storage and transfer of iron. Structure of

**Unit III Ionic equilibrium (9 h)**

Strong, moderate and weak electrolytes, degree of ionization, factors affecting degree of Ionization, ionization constant and ionic product of water. Ionization of weak acids and Bases, pH scale, Buffer solutions- Henderson's equation. Indicators-theories of acid – base Indicators, selection of Indicators, Common ion effect Solubility and solubility product of sparingly soluble salts – applications of solubility product principle.

**Unit IV Chemical Kinetics-I: (9 h)**

The concept of reaction rates. Effect of temperature, pressure, catalyst and other factors on Reaction rates. Order and molecularity of a reaction, Derivation of integrated rate equations For zero, first and second order reactions (similar and different reactants). Half-life of a Reaction. General methods for determination of order of a reaction.

**Unit V Chemical Kinetics-II: (9 h)**

Concept of activation energy and its calculation from Arrhenius equation. Theories of Reaction Rates: Collision theory and Activated Complex theory of bimolecular reactions. Comparison of the two theories (qualitative treatment only). Enzyme catalysis- Specificity, factors affecting enzyme catalysis, Inhibitors and Lock & key Model. Michaels- Menten equation- derivation, significance of Michaelis-Menten constant.

**Reference books**

- 1) Text book of physical chemistry by S Glasstone
- 2) Concise Inorganic Chemistry by J.D.Lee
- 3) Advanced physical chemistry by Gurudeep Raj
- 4) Advanced physical chemistry by Bahl and Tuli
- 5) Inorganic Chemistry by J.E.Huheey
- 6) Basic Inorganic Chemistry by Cotton and Wilkinson.

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## IV - SEMESTER

Physical Chemistry - Volumetric Analysis (Practical)

Credits: 01

Physical Chemistry - Volumetric Analysis

### Course outcomes:

At the end of the course, the student will be able to:

1. Use glassware, equipment and chemicals and follow experimental procedures in the Laboratory
2. Understand and explain the volumetric analysis based on fundamental concepts learnt in Ionic equilibrium
3. Learn and identify the concepts of standard solutions, primary and secondary standards Facilitate the learner to make solutions of various molar concentrations.

### Syllabus:

#### Volumetric analysis:

- Estimation of sodium hydroxide using standardized HCl solution.
- Estimation of sodium carbonate and sodium hydroxide present in a mixture.
- Determination of Fe (II) using  $\text{KMnO}_4$  with oxalic acid as primary standard. (Internal Indicator method)
- Determination of Fe (II) using  $\text{KmnO}_4$  with oxalic acid as primary standard. (External Indicator method)
- Estimation of water of crystallization in Mohr's salt by titrating with  $\text{KmnO}_4$
- VI. Co-curricular activities and assessment methods:
  - Continuous Evaluation: Monitoring the progress of student's learning
  - Class Tests, Worksheets and Quizzes
  - Presentations, Projects and Assignments and Group Discussions: Enhances critical thinking skills and personality

#### List of reference books:

A Text Book of Quantitative Inorganic Analysis (3rd Edition) –A.I.Vogel

Web related references suggested by teacher.

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II B.Sc. Honours (Botany)- Semester-IV  
Minor- Chemistry

Title of the Paper-3 PHYSICAL CHEMISTRY-II  
(STATES OF MATTER, PHASE RULE AND SURFACE CHEMISTRY)  
**MODEL QUESTION PAPER**

Time: 3 Hours

Max. Marks: 70

SECTION- A

Answer any FIVE questions

(5 x 4 = 20 Marks)

1. What is congruent and incongruent melting point? Give examples.
2. Write phase rule and explain the terms phase, Component, degrees of freedom.
3. Derive Boyle's Charles law from kinetic gas equation.
4. Explain Joule-Thomson effect.
5. Discuss the applications of liquid crystals.
6. Explain briefly about the effect of addition of various solutes on the surface tension.
7. Define lattice point, space lattice, unit cell with examples.
8. Explain Frenkel and schottky defects of crystals.
9. Explain Hardy-Schultz rule of colloids.
10. Explain about sol and emulsion.

SECTION- B

(5 x 10 = 50 marks)

Answer ALL the following questions.

11. (a) Discuss the main postulates of kinetic theory of gases.  
(b) Relationship between critical constants and Vanderwaals constants.  
Or
12. Derive Vanderwaals equation.
13. Define coefficient of viscosity of liquids. Explain the determination of coefficient of viscosity of liquids.  
Or
14. (a) Discuss the differences between liquid crystals and solids.  
(b) Discuss the nematic and smectic liquid crystals.
15. Derive Bragg's equation and also discuss powder method.  
Or
16. (a) Explain law of interfacial angle.  
(b) Explain the law of rationality of indices.
17. Explain silver-lead system.  
Or
18. Discuss Salt-water system.
19. What is adsorption? Discuss the differences between physical adsorption and chemical adsorption.  
Or
20. Explain ((a) Stability of colloids and (b) Gold number

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Title: Paper-4-GENERAL AND PHYSICAL CHEMISTRY  
MODEL QUESTION PAPER

Time: 3 Hours

Max. Marks: 70

**SECTION- A**

Answer any FIVE questions

(5 x 4 = 20 Marks)

1. Explain Hund's rule and Aufbau principle with examples.
2. Define ionization potential and explain its periodic trend.
3. What is lattice energy? Explain the factors affecting lattice energy.
4. Explain Fajan's rules and its significance in determining the covalent character in ionic compounds.
5. Define hybridization and describe the geometry of CH<sub>4</sub> and SF<sub>6</sub> molecules.
6. Discuss the importance of hydrogen bonding with suitable examples.
7. Explain Arrhenius and Bronsted-Lowry theories of acids and bases.
8. Define pH and discuss its significance in acid-base chemistry.
9. Explain free electron theory and band theory in metallic bonding.
10. Define lattice energy and discuss its role in ionic bonding.

**SECTION- B**

5 x 10 = 50 marks

Answer ALL the following the questions.

11. Explain the Bohr Theory and its limitations.  
Or
12. Discuss the periodic properties of elements such as atomic radii, ionic radii, and electron affinity.
13. Discuss the properties of ionic compounds and the factors favoring their formation  
Or
14. Explain the concept of polarization and its effects on solubility, melting point, and stability of ionic compounds.
15. Describe the Valence Bond Theory and hybridization of atomic orbitals in BF<sub>3</sub> and PCl<sub>5</sub>.  
Or
16. Construct and explain the Molecular Orbital diagrams for N<sub>2</sub> and CO.
17. Explain the Band Theory of Metals and classify substances as conductors, semiconductors, and insulators.  
Or
18. Discuss the types of weak bonds and their significance in molecular interactions.
19. Explain Pearson's HSAB principle and its importance.  
Or
20. Describe acid-base reactions in non-aqueous solvents like liquid ammonia.

*Y.v. Lakshmi*

**LIST OF THE QUESTION PAPER SETTERS**

SL No	Details of the Expert	Institutional Address	Mobile No.	Mail address
1	DR. S. VENKATA LAKSHMI REDDY	GDCW, RAYACHOTI	9063467494	Svlreddy2003@gmail.com
2	P V RAMANA REDDY	GDC, YERRAGUNTLA	9052074029	Ramanareddypv29@gmail.com
3	S. SOWJANYA	SRI HARI DEGREE COLLEGE, KADAPA	8341471728	Sowjisomireddy@gmail.com
4	M.PAVAN KUMAR REDDY	SKSC DEGREE COLLEGE, PRODDATUR.	8919034475	Moola.pavan19@gmail.com
5	G. MALLIKHARJUN A REDDY	SAI PARAMESWARA DEGREE COLLEGE, JAMMALAMADUGU	9494123613	Gopireddymalli222@gmail.com
6	Dr. B. Anandakumar	Sri YN College(A) Narsapur-534275 West Godavari district	9912020823	Bokkaak2003@gmail.com

**List of Examiners**

1	V. BALA NARASIMHA REDDY	SVCOLLEGE OF HIGHER EDUCATION, MYDYKU R.	9441502226	Vbnreddy009@gmail.com
2	M.MADHAVI	NAGARJUNA DEGREE COLLEGE FOR WOMEN, KADAPA	9573299321	Mallumadhavi95@gmail.com
3	P. Sri Lakshmi	SRINIVASA DEGREE COLLEGE, JMD	80194377545	Sri.pallalakshmi@gmail.com
4	D.SAMBASIVUDU	MEDHA DC, MYDUKUR.	8919945852	Sambasiva222@gmail.com
5	O. RAJASEKHAR	SVCHE, MYDUKUR	8374047579	rajasekharobulapu@gmail.com

*Y.v. Lakshmi*

Department of chemistry  
2025-2026

Members present for the Board of Studies:

I.No.	Name of the Faculty	Designation	Position	Signature
01	Y Venkata Lakshmi	Head of the Department	Chairman	<i>Y.V. Lakshmi</i>
02	Dr..K.S.V.Krishna Rao	Professor	University Nominee	Virtual
03	Dr. O Akbar	Assistant Professor	Subject Expert, Annamacharya University	Virtual
04	Dr.S Gangadhara	Assistant Professor	Subject Expert, RGUKT, Ongole	Virtual
05	T.Venkata Prasad	Industrialist	Member	Virtual
06	P.Manoj Kumar Yadav	Alumni	Member	Virtual

Chemistry BoS Meeting-I: 2024-25

Saturday, June 14 · 10:30am – 12:00pm

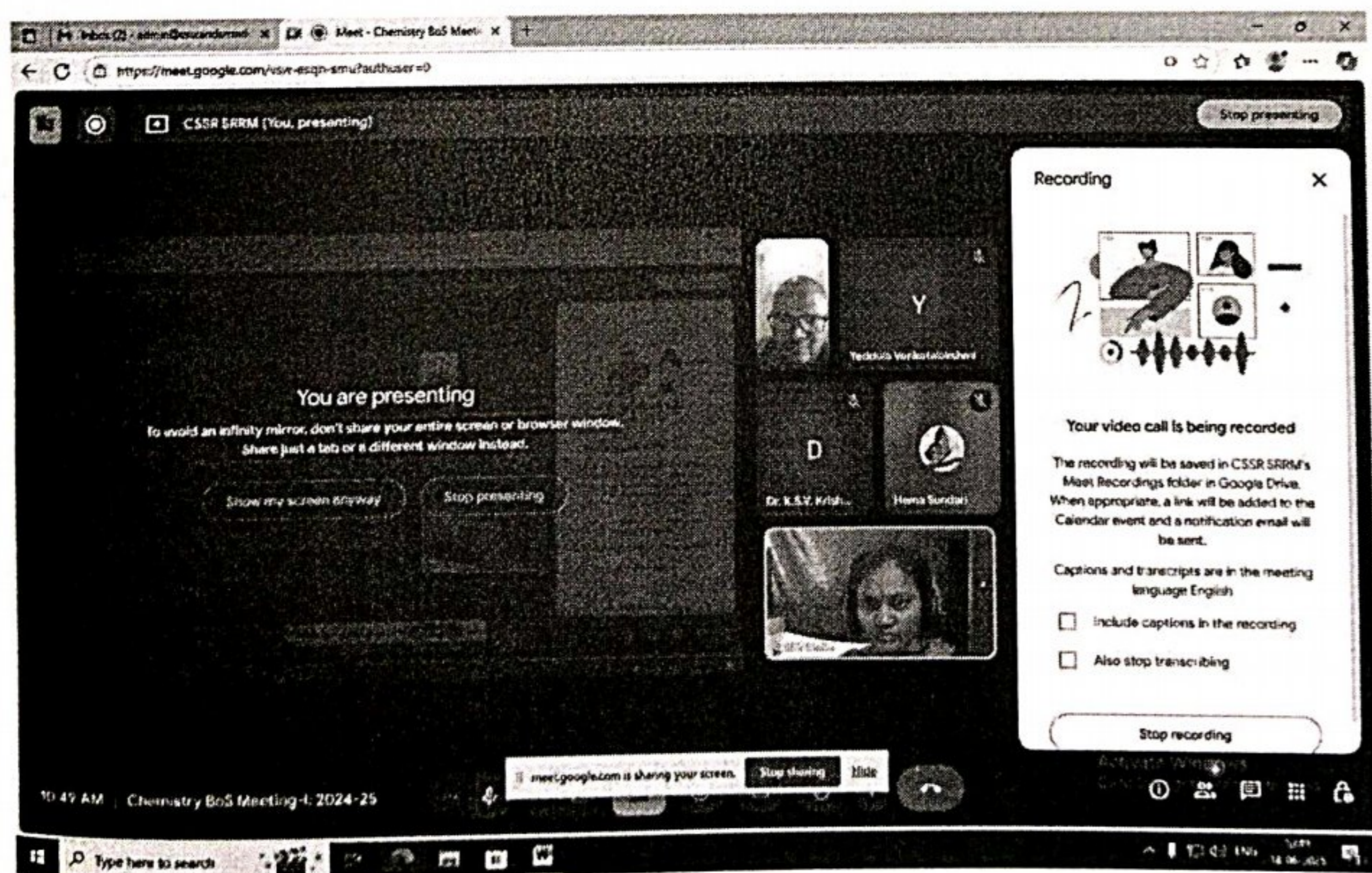
Time zone: Asia/Kolkata

Google Meet joining info

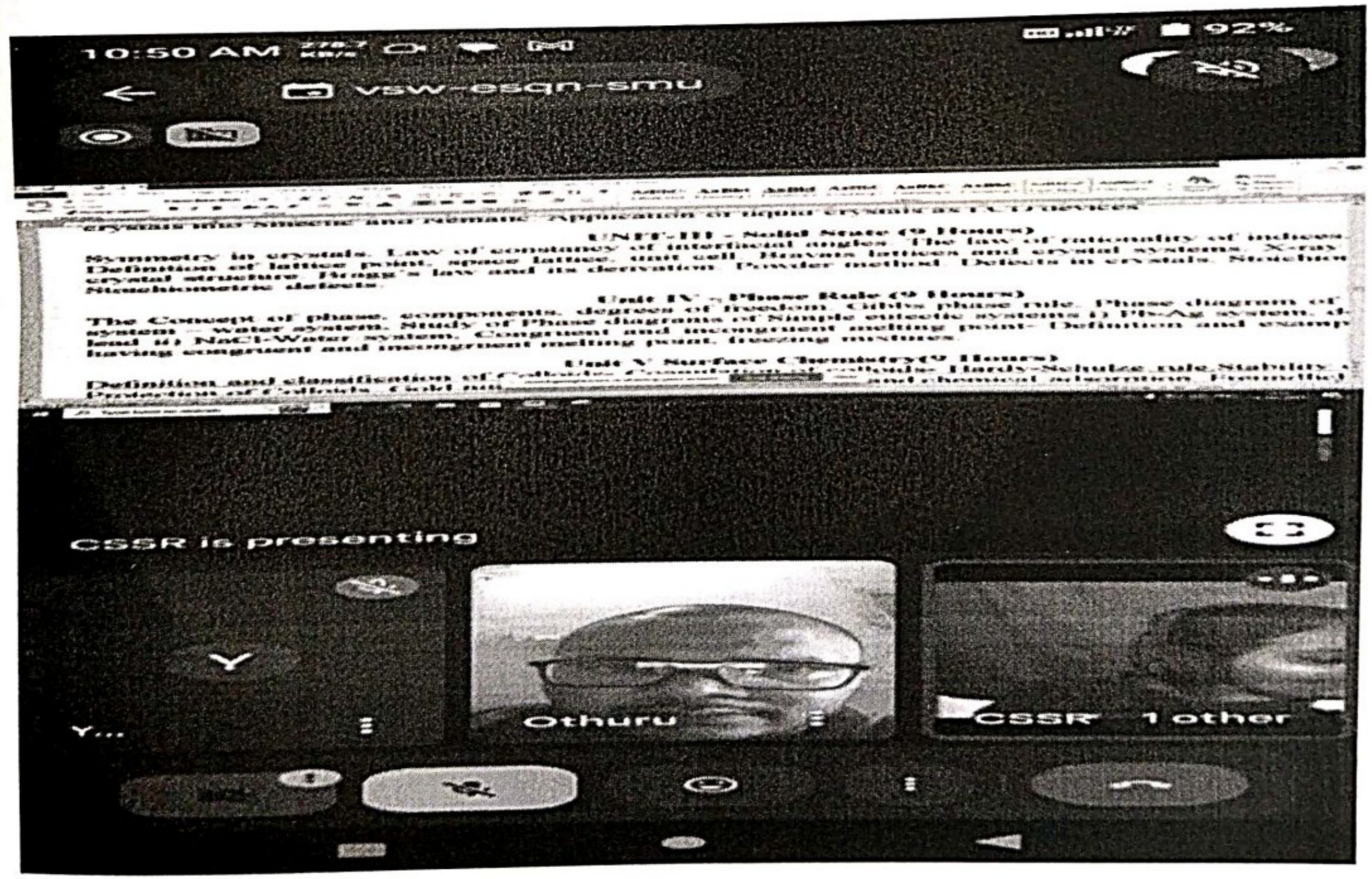
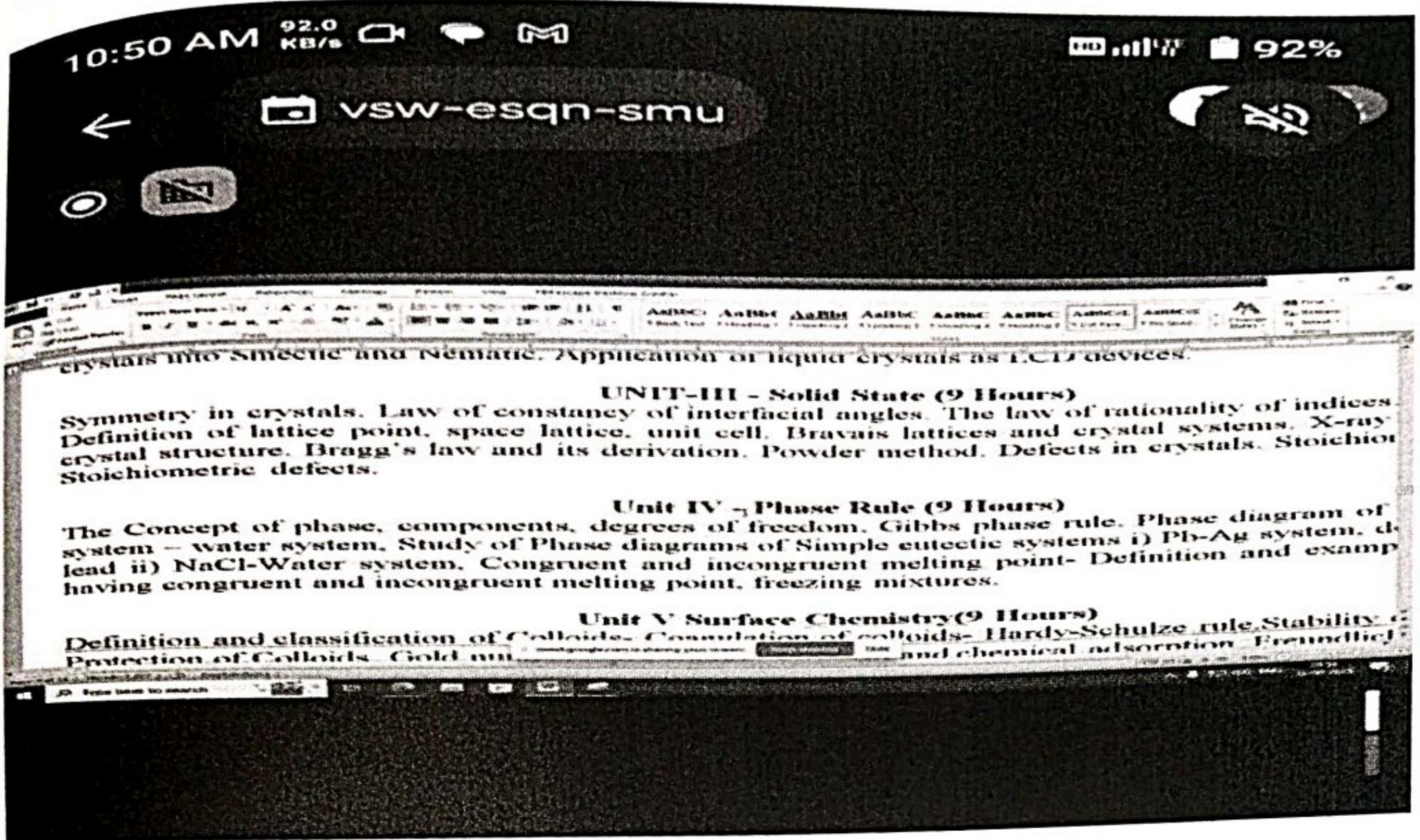
Video call link: <https://meet.google.com/vsw-esqn-smu>

Or dial: (US) +1 651-571-0768 PIN: 624 171 029#

More phone numbers: <https://tel.meet/vsw-esqn-smu?pin=1461710412196>



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